

General Chemistry
Mr. MacGillivray
Quiz #13:
Types of Bonding

This type of bonding ... Ionic Covalent Metallic	... is likely to take place between these types of elements: _____ and _____ _____ and _____ _____ and _____	... and involves this types of electron interaction : The electrons are _____ ed between atoms The electrons are _____ ed between atoms The electrons are _____ ed between atoms
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If electrons are shared **unevenly** between atoms, this is called
_____ bonding.

If electrons are shared **evenly** between atoms, this is called
_____ bonding.

X is an unknown element. If X is a nonmetal, would you expect a molecule with the formula "X₂" to have a polar bond?

Why or why not?

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This type of bonding	... is likely to take place between these types of elements:	... and involves this types of electron interaction :
...		
Ionic	_____ and _____	The electrons are _____ ed between atoms
Covalent	_____ and _____	The electrons are _____ ed between atoms
Metallic	_____ and _____	The electrons are _____ ed between atoms

If electrons are shared unevenly between atoms, this is called

_____ bonding.

If electrons are shared evenly between atoms, this is called

_____ bonding.

X is an unknown element. If X is a nonmetal, would you expect a molecule with the formula "X₂" to have a polar bond?

Why or why not?

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This type of bonding	... is likely to take place between these types of elements:	... and involves this types of electron interaction :
...		
Ionic	<u>met</u> <u>al</u> and <u>nonmet</u> <u>al</u>	The electrons are <u>transf</u> <u>ered</u> between atoms
Covalent	<u>nonmet</u> <u>al</u> and <u>nonmet</u> <u>al</u>	The electrons are <u>shar</u> <u>ed</u> between atoms
Metallic	<u>met</u> <u>al</u> and <u>met</u> <u>al</u>	The electrons are <u>pool</u> <u>ed</u> between atoms

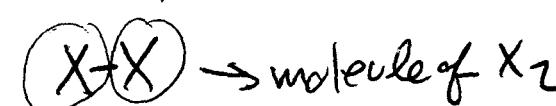
If electrons are shared unevenly between atoms, this is called

polar covalent bonding.

If electrons are shared evenly between atoms, this is called

nonpolar covalent bonding.

X is an unknown element. If X is a nonmetal, would you expect a molecule with the formula "X₂" to have a polar bond?

Why or why not? No whatever X's electronegativity is, the diff in electroneg. between X and X = 0.


Therefore, the bond is nonpolar